

Name \_\_\_\_\_

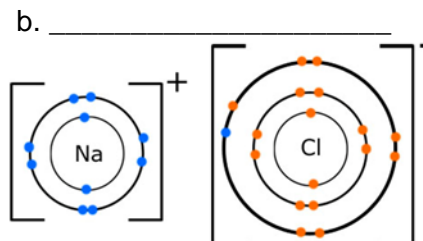
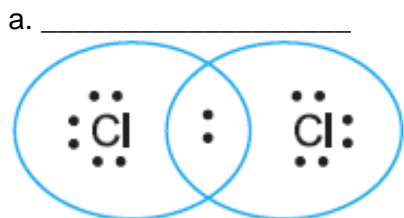
## Phusi 4 Practice 3

Target	1	2 (all of 1 plus)	3 (all of 2 plus)	4 (all of 3 plus)
<b>LE 5.7 Preparedness</b>	Does not complete formative or summative in an effortful and timely manner, is not engaged, does not arrive on time with class materials ready to learn, does not communicate when issues arise	Completes formative or summative in an effortful or timely manner, is sometimes engaged, sometimes arrives on time with class materials ready to learn, sometimes communicates when issues arise	Completes formative or summative in an effortful and timely manner, remains engaged, arrives on time with materials ready to learn, communicates when issues arise	Completes formative or summative in an effortful and timely manner, remains engaged, arrives on time with materials ready to learn, communicates when issues arise, and is reflective on strengths and challenges within your preparedness skill
<b>LE 5.6 Precision</b>	Recognizes the importance of products that are planned, edited, and completed with care	Attempts products that are planned, edited, and completed with care	Creates products that are planned, edited, and completed with minimal errors	Creates products that are planned, edited, and completed free from errors or need for revision
<b>Phusikos 4</b>	I can recognize a chemical bond	I can identify covalent bonding and ionic bonding.	I can <b>draw, build, and/or model</b> covalent bonding and ionic bonding.	I can <b>compare</b> and <b>contrast</b> covalent bonding and ionic bonding.
<b>MP2 Atoms, Bonding</b>	I can diagram the shell structure of an atom and an understanding of valence electrons	(all of 1 plus) I can use the periodic table to predict properties of atoms of elements based on patterns of electrons in atoms	(all of 2 plus) I can predict and diagram bonding between atoms	(all of 3 plus) Nailed it!

1. Draw a **Lewis Dot Structure** for the following atoms.

Oxygen	Chlorine	Magnesium

2. Identify the **types of bonds** represented in the images below



3. Compare and contrast covalent and ionic bonding.

a. How are the types of bonds similar?

b. How are the types of bonds different?

c. Atoms of which type(s) (elements) form ionic bonds and which form covalent bonds?

4. Draw Lewis Dot Diagrams to represent the bonding for each of the element pairs.

Represent enough atoms of each element to account for enough electrons getting transferred between atoms to get FOS, or enough electrons getting shared between atoms to FOS.

Lewis Dot Diagram of atom (A)	Lewis Dot Diagram of atom (B)	Covalent or Ionic Bond?	Lewis Dot Diagram of the bonded atoms		
			Atom(s) A	Atom(s) B	Bonded atoms
1. <p style="text-align: center;"><b>H</b></p> 2. Metal or nonmetal? 6. Gain / lose / share $e^-$ 7. # $e^-$ ____	3. <p style="text-align: center;"><b>Cl</b></p> 4. Metal or nonmetal? 6. Gain / lose / share $e^-$ 7. # $e^-$ ____	5.			
1. <p style="text-align: center;"><b>Li</b></p> 2. Metal or nonmetal? 6. Gain / lose / share $e^-$ 7. # $e^-$ ____	3. <p style="text-align: center;"><b>F</b></p> 4. Metal or nonmetal? 6. Gain / lose / share $e^-$ 7. # $e^-$ ____				
1. <p style="text-align: center;"><b>N</b></p> 2. Metal or nonmetal? 6. Gain / lose / share $e^-$ 7. # $e^-$ ____	3. <p style="text-align: center;"><b>H</b></p> 4. Metal or nonmetal? 6. Gain / lose / share $e^-$ 7. # $e^-$ ____				
1. <p style="text-align: center;"><b>C</b></p> 2. Metal or nonmetal? 6. Gain / lose / share $e^-$ 7. # $e^-$ ____	3. <p style="text-align: center;"><b>O</b></p> 4. Metal or nonmetal? 6. Gain / lose / share $e^-$ 7. # $e^-$ ____				