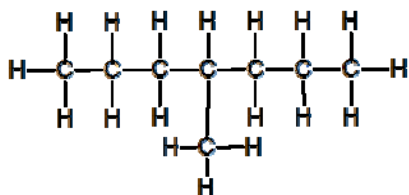


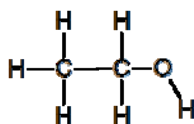
Polarity of BIG Molecules and Ionic Substances, and one small molecule

Target	1	2 (all of 1 plus)	3 (all of 2 plus)	4 (all of 3 plus)
LE 5.7 Preparedness	Does not complete formative or summative in an effortful and timely manner, is not engaged, does not arrive on time with class materials ready to learn, does not communicate when issues arise	Completes formative or summative in an effortful or timely manner, is sometimes engaged, sometimes arrives on time with class materials ready to learn, sometimes communicates when issues arise	Completes formative or summative in an effortful and timely manner, remains engaged, arrives on time with materials ready to learn, communicates when issues arise	Completes formative or summative in an effortful and timely manner, remains engaged, arrives on time with materials ready to learn, communicates when issues arise, and is reflective on strengths and challenges within your preparedness skill
LE 5.6 Precision	Recognizes the importance of products that are planned, edited, and completed with care	Attempts products that are planned, edited, and completed with care	Creates products that are planned, edited, and completed with minimal errors	Creates products that are planned, edited, and completed free from errors or need for revision
Hydro 5	I can draw Lewis structures of molecules and formula units.	I can determine shapes of molecules.	I can use electronegativities and geometries to predict polarity.	
MP3 Hydrosphere (Shape, Polarity, Solubility)	I can create a Lewis structure of a particle of a substance, and use it to predict the shape of a molecule.	(all of 1 plus) I can use electronegativities of atoms and shapes of molecules to predict polarity	(all of 2 plus) I can demonstrate an understanding of the relationship between polarity and solubility	(all of 3 plus) and it



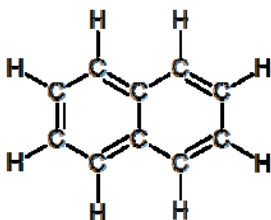
fuel (hydrocarbons)

Similar to _____ (in notes)
 Polar / Not Polar / Both _____
 Explain _____



ethanol

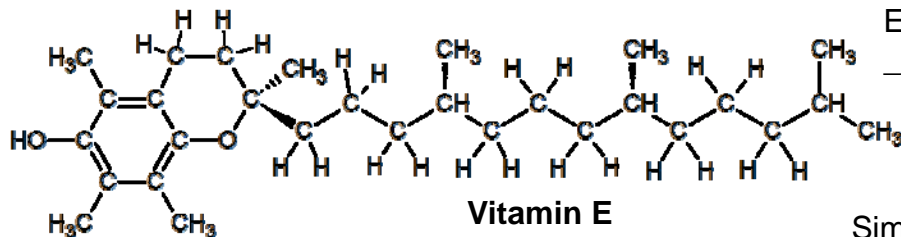
Similar to _____ (in notes)
 Polar / Not Polar / Both _____
 Explain _____



naphthalene

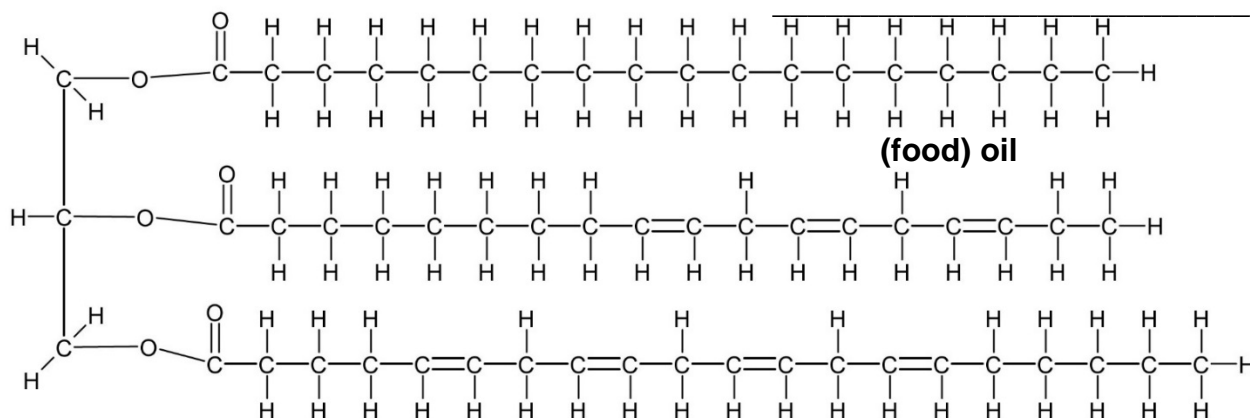
Similar to _____ (in notes)
 Polar / Not Polar / Both _____
 Explain _____

Similar to _____ (in notes)
 Polar / Not Polar / Both _____
 Explain _____

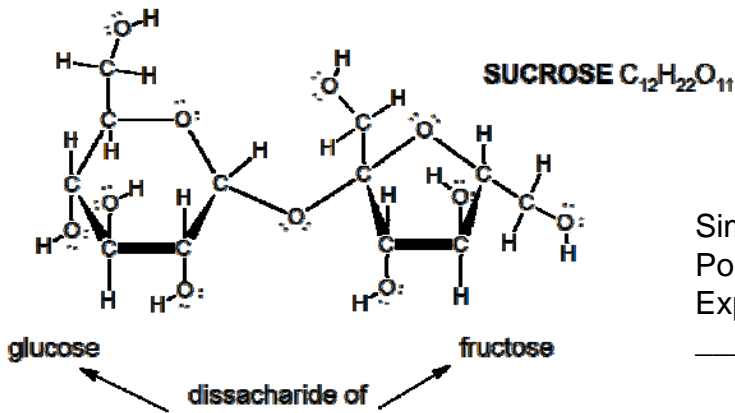


Vitamin E

Similar to _____ (in notes)
 Polar / Not Polar / Both _____
 Explain _____



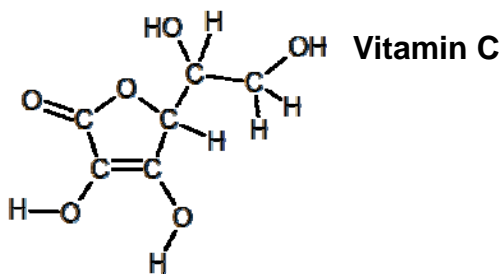
(food) oil



Similar to _____ (in notes)

Polar / Not Polar / Both _____

Explain _____



Similar to _____ (in notes)

Polar / Not Polar / Both _____

Explain _____

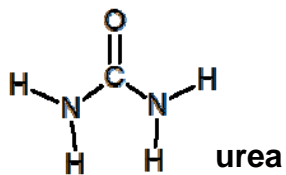


Similar to _____ (in notes)

Polar / Not Polar / Both _____

Explain _____

small-ish molecule
use small molecule rules
CA: C



Similar to _____ (in notes)

Polar / Not Polar / Both _____

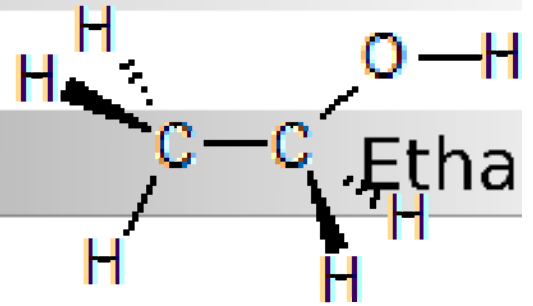
Explain _____

the more like the solvent and substance
the more likely the substance will dissolve

C_xH_y

Hydrocarbons

most soluble

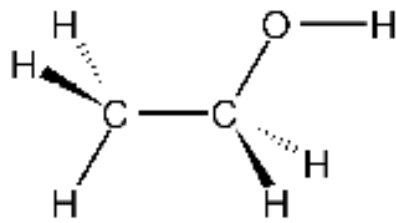


Water

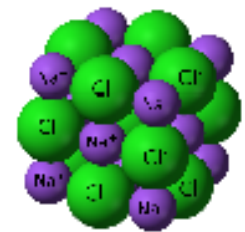
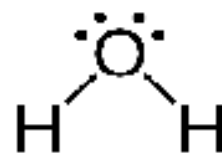


most soluble

carbons
(H_4)



Ethanol



Polar

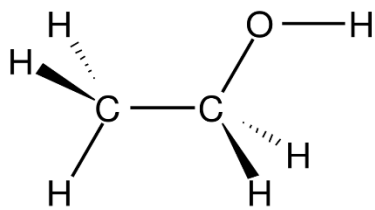
Ionic

Hydrocarbons

Ethano

Water

carbons
(H₄)



Ethanol

